

Chapter 9 Chemical Names Formulas Section Review Answer Key

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Chapter 9: Chemical Names and Formulas. STUDY. PLAY. Monatomic Ion. Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence electrons and behaves as a unit, yet it still has a charge. (+ or -). Determined by the amount of electrons lost or gained.

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To name a binary molecular compound, use the following guidelines: 1. Write the names of the elements in the order listed in the formula 2. Use prefixes appropriately to indicate the number of each kind of atom 3. End the name of the second element with the suffix -ide

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First, confirm that the compound is a binary molecular compound. Name the elements in the order listed in the formula. Use prefixes to indicate the number of each kind of atom. Omit the prefix -mono when the formula contains only one atom of the first element in the name. The suffix of the name of the second element is -ide.

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1. when the name of the anion (X) ends in -ide the acid name begins with the prefix hydro-. The stem of the anion has the suffix -ic and is followed by the word acid (ex. HCl= hydrochloric acid) 2. when the anion name ends in -ite the acid name is the stem of the anion with the suffix -ous followed by the word acid (ex. H₂SO₃ -- X=sulfite = sulfurous acid)

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Writing the Names and Formulas for Bases. A base is a compound that produces OH⁻ in water. When naming a base, you name it like any other compound that starts with a regular or transition metal. Ex: NaOH = sodium hydroxide. When writing the formula for a base, remember to balance charges. Ex: magnesium hydroxide = Mg(OH)₂

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Chapter 9 Chemical Names and Formulas 83 SECTION 9.3 NAMING AND WRITING FORMULAS FOR MOLECULAR COMPOUNDS (pages 268–270) This section explains the rules for naming and writing formulas for binary molecular compounds. Naming Binary Molecular Compounds (pages 268–269) 1. Circle the letter of the type(s) of elements that form binary molecular compounds.

~~Name Date Class CHEMICAL NAMES AND FORMULAS 9~~

SECTION 9.2 NAMING AND WRITING FORMULAS FOR IONIC COMPOUNDS I. Write the formulas for these binary ionic compounds. c. potassium iodide 2. write the formulas for the compounds formed c. b. f. 3. Name the following binary ionic compounds. sodium sulfide g. CuCl: h. SnCl. a. MnO₂ c. d. SrBr₂ 221 Chapter 9 Chemical Names and Formulas

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Chemical Names and Formulas 281 CHAPTER 9 Assessment 42. a. 2- b. 1+ c. 1- d. 3+ 43. a. 2+ b. 2+ c. 3+ d. 1+ 44. a. barium ion b. iodide ion c. silver ion d. mercury(II) ion 45. cyanide, CN⁻ and hydroxide, OH⁻ 46. a. hydroxide ion b. lead(IV) ion c. sulfate ion d. oxide ion 47. zero; A compound is electrically neutral. 48. The symbols for the cation and

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Chapter 9 Chemical Names Formulas Chapter 9 - Chemical Names and Formulas. STUDY. Flashcards. Learn. Write. Spell. Test. PLAY. Match. Gravity. Created by. bibiloni1901. Terms in this set (20) Monatomic Ion. Consists of a single atom with a positive or negative charge resulting from the loss of gain of one or more valence

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Chapter 9 "Chemical Names and Formulas" Tools. Copy this to my account; E-mail to a friend; Find other activities; Start over; Help; Use these activities to help yourself to learn the vocabulary and major topics presented in this chapter. A B; anion: any atom or group of atoms that has a negative charge:

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Chapter 9 Practice Test - Naming and Writing Chemical Formulas Answer Section MATCHING 1. ANS: A PTS: 1 DIF: L1 REF: p. 253 OBJ: 9.1.1 Identify the charges of monatomic ions by using the periodic table, and name the ions. STA: Ch.5.a 2.ANS: H PTS: 1 DIF: L1 REF: p. 254

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Chemistry: Imagination and Implication focuses on the importance and impact of chemistry on daily living. This book discusses the essential concepts of chemistry and its application. Organized into 16 chapters, this book starts with an overview of the experimental facts, principles, and methods of chemistry as an aid in exercising intelligent and informed judgment in instances where controversy surrounds the interaction of chemistry with society or the individual. This text then explores the practical arts of metallurgy, which achieved a considerable degree of sophistication long before they were scientifically understood. The reader is then introduced to the atomic concept, the conservation of mass, as well as to the substances that constitute the living things. Other chapters consider the polymerization of amino acids into peptides and proteins. The final chapter examines the various applications of radioactive isotopes produced in particle accelerators. This book is intended for students and teachers who are involved in a chemistry course.

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